

## Isolated spin 3/2 plaquettes in $\text{Na}_3\text{RuO}_4$

K.A. Regan<sup>a,\*</sup>, Q. Huang<sup>b</sup>, R.J. Cava<sup>a</sup>

<sup>a</sup>Department of Chemistry, Princeton University, Princeton, NJ 08544, USA

<sup>b</sup>NIST Center for Neutron Research, National Institute of Standards and Technology, Gaithersburg, MD 20899, USA

Received 16 February 2005; received in revised form 6 April 2005; accepted 12 April 2005

### Abstract

The crystal structure of  $\text{Na}_3\text{RuO}_4$  determined by powder neutron diffraction is reported. The structure consists of isolated tetramers of edge sharing  $\text{RuO}_6$  octahedra in the  $ab$  plane, creating isolated four-member plaquettes of Ru atoms comprised of two equilateral triangles sharing an edge. Magnetic susceptibility measurements reveal an antiferromagnetic transition at  $\sim 29$  K, with  $\theta_w = -141$  K. Neutron diffraction data indicate the onset of three-dimensional magnetic ordering at 29 K.

© 2005 Elsevier Inc. All rights reserved.

**Keywords:** Ruthenates;  $\text{Na}_3\text{RuO}_4$ ; Magnetic properties.

### 1. Introduction

Ruthenates are rich in structure and varied in properties. Perovskite-based alkaline earth ruthenates ( $M\text{RuO}_3$  or  $M_{n+1}\text{Ru}_n\text{O}_{2n+1}$ , where  $M = \text{Ca}, \text{Sr}, \text{Ba}$ ), exhibit properties varying from ferromagnetism in  $\text{SrRuO}_3$  [1], to weakly temperature dependent Pauli paramagnetism in  $\text{BaRuO}_3$  [2] to superconductivity in  $\text{Sr}_2\text{RuO}_4$  [3]. Characterization of  $\text{Sr}_3\text{Ru}_2\text{O}_7$  at higher temperatures point towards ferromagnetic interactions between the Ru atoms, while at lower temperatures, the compound crosses over to antiferromagnetic-like behavior [4] and becomes metamagnetic at high applied fields [5].  $\text{CaRuO}_3$  can be viewed in the same manner: its paramagnetic behavior can be tipped toward ferromagnetism by doping [6,7]. Such varied behavior of properties suggests that ruthenates sit on the boundary between magnetic and non-magnetic states.

The alkali metal ruthenates have not been as extensively studied. Recent work on the hollandites,  $\text{KRu}_4\text{O}_8$  and  $\text{RbRu}_4\text{O}_8$ , revealed temperature independent Pauli paramagnetism [8] and metallic conductivity. Pauli paramag-

netism has also been observed in  $\text{Li}_2\text{RuO}_3$  [9]. Although the magnetic properties of these three compounds are relatively ordinary, a further investigation of the alkali metal ruthenates may expose compounds that exhibit more unusual properties. For instance,  $\text{Li}_3\text{RuO}_4$  exhibits short range antiferromagnetic ordering, with the possibility of spin freezing at  $\sim 10$  K [10]. The Na–Ru–O system, for example, has not been thoroughly investigated, with several previously reported phases only superficially characterized [11]. Recently, Darriet et al. reported the structure and the magnetic properties of  $\text{NaRuO}_2$  as well as  $\text{Na}_2\text{RuO}_4$  [12,13].  $\text{NaRuO}_2$  displays paramagnetic behavior, while  $\text{Na}_2\text{RuO}_4$  exhibits antiferromagnetic behavior attributed to short range ordering.

Although  $\text{Na}_3\text{RuO}_4$  has been reported as being analogous in structure to  $\text{Na}_3\text{NbO}_4$ , the crystal structure was not previously refined, and magnetic data had only been collected from 77 to 600 K [14]. Mössbauer  $^{99}\text{Ru}$  data were interpreted in terms of an antiferromagnetic long range ordering in  $\text{Na}_3\text{RuO}_4$  at low temperature [15]. In this paper, we present the neutron diffraction refinement of  $\text{Na}_3\text{RuO}_4$  and low temperature magnetic characterization. It is postulated that long range ordering of the spin 3/2  $\text{Ru}^{5+}$  magnetic moments is frustrated by the isolated equilateral plaquettes present in the structure.

\*Corresponding author. Fax: +609 258 6746.

E-mail address: [kregan@princeton.edu](mailto:kregan@princeton.edu) (K.A. Regan).

## 2. Experimental

Polycrystalline samples of  $\text{Na}_3\text{RuO}_4$  were synthesized using stoichiometric amounts of NaOH pellets (Merck, 97.0%) and  $\text{RuO}_2$  powder (Alfa Aesar, Ru 54% min). The  $\text{RuO}_2$  powder was heated prior to sample preparation at 700 °C for 2 h to remove any absorbed water. Reaction mixtures were placed in dense alumina crucibles and heated first at 500 °C under flowing  $\text{O}_2$  for 18 h and then at 650 °C under  $\text{N}_2$  for 18 h, with an intermediate grinding.

Phase purity was determined via powder X-ray diffraction using  $\text{CuK}\alpha$  radiation. Resulting patterns corresponded with previously reported peak positions [14]. Magnetic characterization was performed using a Quantum Design PPMS magnetometer. The neutron powder diffraction intensity data of  $\text{Na}_3\text{RuO}_4$  were collected at the NIST Center for Neutron Research, on a high resolution powder neutron diffractometer, with monochromatic neutrons of wavelength 1.5403 Å produced by a Cu(311) monochromator. Collimators with horizontal divergences of 15', 20' and 7' of arc were used before and after the monochromator and after the sample, respectively. Data were collected in the  $2\theta$  range of 3° and 168°, with a step size of 0.05°. The structural parameters were refined using the program GSAS [16]. The neutron scattering amplitudes used in the refinement were 0.363, 0.721, and 0.581 ( $10^{-12}$  cm) for Na, Ru and O, respectively. The magnetic peak intensity at  $2\theta = 23.1^\circ$  as a function of temperature for  $\text{Na}_3\text{RuO}_4$  was measured with neutrons of wavelength 2.359 Å on the BT-2 triple-axis spectrometer at the NIST Center for Neutron Research, with a PG filter employed to suppress higher order wavelength contaminations.

## 3. Results

The structure of  $\text{Na}_3\text{RuO}_4$  was refined to a monoclinic space group  $C2/m$  (#12) with lattice parameters of  $a = 11.0295(6)$  Å,  $b = 12.8205(7)$  Å,  $c = 5.7028(3)$  Å, and  $\beta = 109.90(3)^\circ$ , and the initial atomic positions were taken as those of  $\text{Na}_3\text{TaO}_4$  [14]. The neutron diffraction pattern is shown in Fig. 1. Atomic positions and thermal parameters are listed in Table 1: the final structure was very close to that of  $\text{Na}_3\text{TaO}_4$ . Relevant Ru–O and Na–O bond distances, as well as Ru–O bond angles, are listed in Tables 2 and 3, respectively. The average Ru–O bond length ( $= 2.01$  Å) is similar to that reported for other  $\text{Ru}^{5+}$  compounds [17–19]. The oxygen coordination to the Ru atom creates two distinct octahedral environments (see Fig. 3). The octahedra are distorted both in bond lengths and bond angles. The distortions are similar to those seen in other ruthenates, such as  $\beta\text{-La}_3\text{RuO}_7$  and  $\text{La}_2\text{Sr}_2\text{LiRuO}_8$  [20,21]. In the present case, the irregular Ru–O bond lengths in the  $\text{RuO}_6$

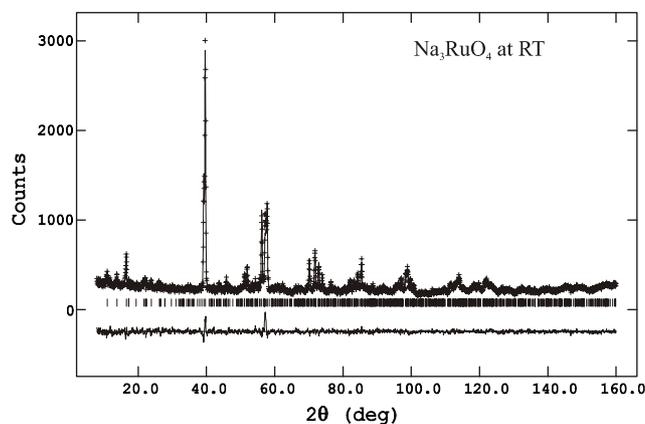


Fig. 1. Observed intensities (crosses) and calculated neutron diffraction pattern (solid line)  $\text{Na}_3\text{RuO}_4$  at 295 K. Vertical lines show reflection positions. Differences between the observed and calculated intensities are shown at the bottom of the figure.

Table 1

Atomic positions and thermal parameters of  $\text{Na}_3\text{RuO}_4$ , e.g.,  $C2/m$  (#12)  $a = 11.0295(6)$  Å,  $b = 12.8205(7)$  Å,  $c = 5.7028(3)$ ,  $\beta = 109.908(3)$ ,  $Z = 4$

| Atom | Site | $x$           | $y$           | $z$         | $U_{iso} \times 100$ (Å <sup>2</sup> ) |
|------|------|---------------|---------------|-------------|--|
| Ru1  | 4g   | 0             | 0.1252(7)     | 0           | 0.59(11)                               |
| Ru2  | 4i   | 0.2550(9)     | 0             | 0.01777(16) | 0.59(11)                               |
| Na1  | 4g   | 0             | 0.3759(15)    | 0           | 1.28(13)                               |
| Na2  | 4e   | $\frac{1}{4}$ | $\frac{1}{4}$ | 0           | 1.28(13)                               |
| Na3  | 8j   | 0.2421(13)    | 0.1254(10)    | 0.          | 1.28(13)                               |
| Na4  | 2c   | 0             | 0             | 1/2         | 1.28(13)                               |
| Na5  | 4h   | 0             | 0.2565(14)    | 1/2         | 1.28(13)                               |
| Na6  | 2d   | 0             | $\frac{1}{2}$ | 1/2         | 1.28(13)                               |
| O1   | 4i   | 0.1147(14)    | 0             | 0.2019(27)  | 1.07(6)                                |
| O2   | 8j   | 0.1032(8)     | 0.2278(6)     | 0.2285(17)  | 1.07(6)                                |
| O3   | 4i   | 0.1528(12)    | 0             | 0.1969(23)  | 1.07(6)                                |
| O4   | 8j   | 0.1197(9)     | 0.1119(7)     | 0.8111(17)  | 1.07(6)                                |
| O5   | 8j   | 0.1455(8)     | 0.3924(7)     | 0.7636(16)  | 1.07(6)                                |

$R_p = 5.38$ ,  $R_w = 6.45$ ,  $\chi^2 = 1.078$ .

\*Temperature parameters for Ru, Na and O were constrained to be equal, respectively.

octahedra result in equilateral Ru plaquettes, suggesting that Ru–Ru repulsion may play a role in the overall Ru–O geometry.

In a previous publication [14], the structure of  $\text{Na}_3\text{MO}_4$  ( $M = \text{Nb}$ , Ru, and Ta) was viewed as a derivative of the NaCl structure with ordering of the  $\text{NaO}_6$  and  $\text{RuO}_6$  octahedra in planes parallel to the  $b$ -axis. However, from the magnetic point of view, it is advantageous to look at the structure with respect to the coordination and packing of the  $\text{RuO}_6$  octahedra only, as it is the Ru that are responsible for the observed magnetic behavior. By doing so, the structure of  $\text{Na}_3\text{RuO}_4$  can be viewed as consisting of isolated tetramers of edge-sharing  $\text{RuO}_6$  that lie in the  $ab$  plane. The planes of  $\text{RuO}_6$  octahedra are separated by a layer



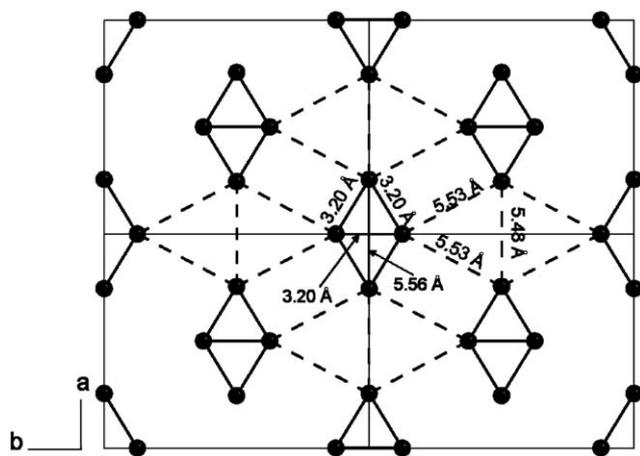


Fig. 4. Structure of  $\text{Na}_3\text{RuO}_4$  in the  $ab$  plane showing the Ru atoms only (4 unit cells). The isolated plaquettes of Ru atoms are shown by bold lines. The distances within the plaquette, 3.2 Å, are much shorter than the distances to the next plaquette. The triangular distribution of neighboring plaquettes is shown by dashed lines.

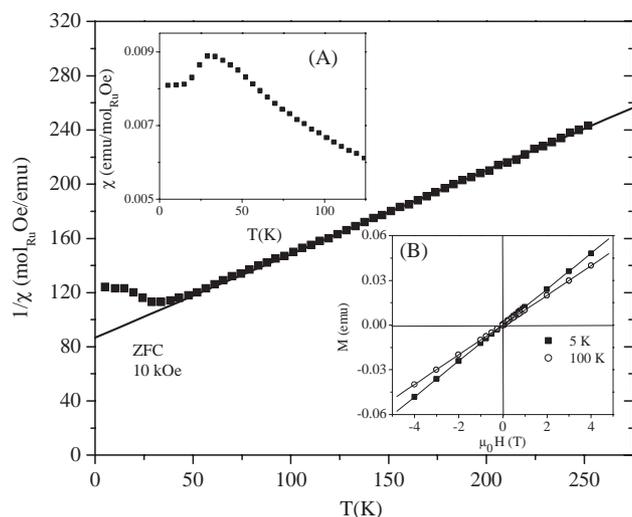


Fig. 5. Inverse magnetic susceptibility ( $1/\chi$ ) vs.  $T$  of  $\text{Na}_3\text{RuO}_4$  with an applied DC field of 10 kOe. A linear fit of the high temperature data (100–250 K) shows Curie–Weiss behavior. Inset A: The antiferromagnetic-like transition observed in  $\text{Na}_3\text{RuO}_4$ . Inset B: The  $M$  vs.  $H$  (0–5 T) loops at 5 and 100 K show linear behavior.

susceptibility from 100 to 250 K (Fig. 5) allowed for the determination of  $\theta_w$  and  $\mu_{\text{eff}}$ . The observed value for  $\mu_{\text{eff}}$  is  $3.60 \mu_B$ . This is similar to the value expected for the spin only moment of an  $S = 3/2$  ion ( $3.87 \mu_B$ ), which is what is expected for  $4d^3 \text{Ru}^{5+}$ . The  $\theta_w$  value was determined to be  $-141$  K. However, temperatures much lower than  $\theta_w$  are needed for the magnetic transition to occur:  $\text{Na}_3\text{RuO}_4$  displays an observed  $\theta_w/T_N$  ratio of approximately 5/1.  $M$  vs.  $H$  loops reveal linear behavior at 5 and 100 K, as would be expected for the antiferromagnetically aligned Ru moments. No ferromagnetic component is observed (Fig. 5, inset B). The  $1/\chi$  data begin to deviate from the high temperature Curie–Weiss

law at  $\sim 55$  K, suggesting the presence of antiferromagnetic fluctuations approximately 25 K above the ordering temperature.

Neutron diffraction data were collected at 3.5 and 48 K on the BT2 diffractometer at NIST. Fig. 6 shows the difference between the diffracted intensity at 3.5 and 48 K. The data show the presence of magnetic ordering at low temperatures, with peaks at  $21.9^\circ$  and  $23.1^\circ$ . The magnetic peak width is that of the instrumental resolution, indicating the ordering is three-dimensional in character. Further, the scattering occurs at positions that not associated with Bragg peaks, verifying that the ordering is antiferromagnetic in character. Fig. 7 plots the intensity of the strongest magnetic scattering ( $2\theta = 23.1^\circ$ ) as a function of temperature, on both heating and cooling. The onset of the 3-D ordering occurs at the same temperature as the main feature in the magnetic susceptibility data. The neutron diffraction data are consistent with the analysis of Mossbauer effect data on the compound taken by Gibb et al. [15].

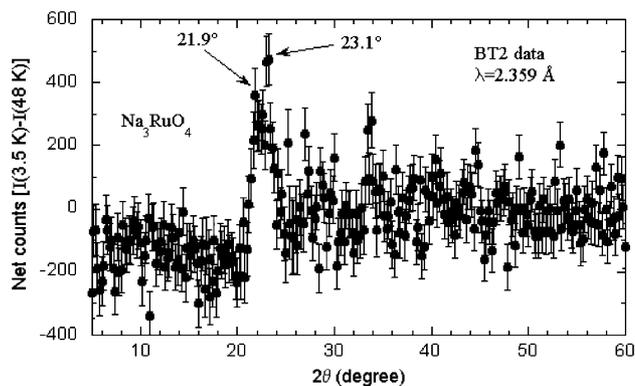


Fig. 6. Difference between the 3.5 and 48 K of the low angle neutron powder diffraction pattern of  $\text{Na}_3\text{RuO}_4$ , in the region in which the strongest magnetic ordering peaks are expected to occur. Magnetic scattering peaks are observed at  $21.9^\circ$  and  $23.1^\circ$ , indicating the presence of magnetic ordering. The position of the peaks (off Bragg peaks) is indicative of antiferromagnetic ordering.

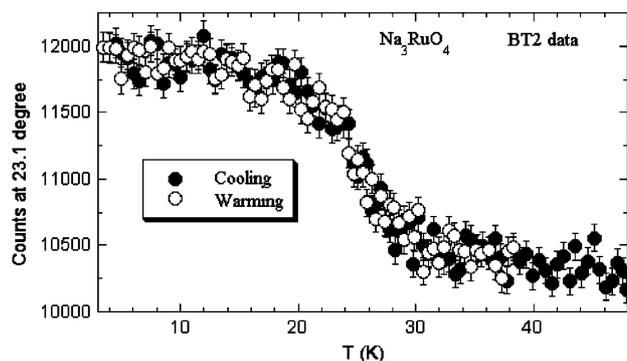


Fig. 7. Temperature dependence of magnetic scattering intensity at  $23.1^\circ$  upon warming and cooling. The onset of magnetic ordering at  $\sim 30$  K corresponds to the ordering temperature observed in the magnetic susceptibility data.

#### 4. Conclusions

We have reported the structure of  $\text{Na}_3\text{RuO}_4$ , determined via powder neutron diffraction. This layered structure is comprised of tetramers of  $\text{RuO}_6$  octahedra, resulting in two equilateral triangles of Ru atoms sharing a single edge to form isolated equilateral plaquettes. Magnetic susceptibility measurements reveal antiferromagnetic ordering of spin  $3/2$  moments at  $T_c \sim 29$  K, approximately a factor of five below  $\theta_w$ , determined to be  $-141$  K. Although such a ratio cannot be considered to be indicative of strong geometric frustration [22], it is possible that the triangular geometry of the Ru atoms suppresses long range ordering observed by neutron diffraction to lower temperatures. We are not aware of theoretical models of systems of isolated plaquettes of this type, and both theoretical modeling and further characterization of the magnetic transition may be of interest.

#### Acknowledgments

This work was supported by the National Science Foundation, grant DMR-0244254.

#### References

- [1] A. Callaghan, C.W. Moeller, R. Ward, *Inorg. Chem.* 5 (1966) 1572.
- [2] J.T. Rijssenbeek, R. Jin, Y. Zadorozhny, Y. Liu, B. Batlogg, R.J. Cava, *Phys. Rev. B* 59 (1999) 4561.
- [3] Y. Maeno, H. Hashimoto, K. Yoshida, S. Nishizaki, T. Fujita, J.G. Bernorz, F. Lichtenberg, *Nature* 372 (1994) 532.
- [4] R.J. Cava, *J. Chem. Soc. Dalton Trans.* 19 (2004) 2979.
- [5] R.S. Perry, L.M. Galvin, S.A. Grigera, L. Capogna, A.J. Schofield, A.P. Mackenzie, M. Chiao, S.R. Julian, S.I. Ikeda, S. Nakatsuji, Y. Maeno, C. Pfleiderer, *Phys. Rev. Lett.* 86 (2001) 2661.
- [6] T. He, R.J. Cava, *Phys. Rev. B* 63 (2001) Art. No. 172,403.
- [7] T. He, R.J. Cava, *J. Phys.: Condens. Matter* 13 (2001) 8347.
- [8] M.L. Foo, W.L. Lee, T. Siegrist, G. Lawes, A.P. Ramirez, N.P. Ong, R.J. Cava, *Mater. Res. Bull.* 39 (2004) 1663.
- [9] I. Felner, I.M. Bradaric, *Physica B* 311 (2002) 195.
- [10] A. Alexander, P.D. Battle, J.C. Burley, D.J. Gallon, C.P. Grey, S.H. Kim, *J. Mater. Chem.* 13 (2003) 2612.
- [11] I.S. Shaplygin, V.B. Lazarev, *Russ. J. Inorg. Chem.* 25 (1980) 1837.
- [12] M. Shikano, C. Delmas, J. Darriet, *Inorg. Chem.* 43 (2004) 1214.
- [13] M. Shikano, R.K. Kremer, M. Ahrens, H.-J. Koo, M.-H. Whangbo, J. Darriet, *Inorg. Chem.* 43 (2004) 5.
- [14] J. Darriet, J. Galy, *Bull. Soc. fr. Minéral. Cristallogr.* 97 (1974) 3.
- [15] T.C. Gibb, R. Greatrex, N.N. Greenwood, *J. Solid State Chem.* 31 (1980) 153.
- [16] A. Larson, R.B. Von Dreele, Los Alamos National Laboratory, Internal Report, 1994.
- [17] J. Darriet, F. Grasset, P.D. Battle, *Mater. Res. Bull.* 32 (1997) 139.
- [18] W.R. Gemmill, M.D. Smith, H.C. zur Loye, *J. Solid State* 177 (2004) 3560.
- [19] E. Quarez, F. Abraham, O. Mentre, *J. Solid State* 176 (2003) 137.
- [20] P. Khalifah, D.M. Ho, Q. Huang, R.J. Cava, *J. Solid State* 165 (2002) 359.
- [21] J.A. Rodgers, P.D. Battle, N. Dupré, C.P. Grey, Sloan, *J. Chem. Mater.* 16 (2004) 4257.
- [22] A.P. Ramirez, *Annu. Rev. Mater. Sci.* 24 (1994) 453.